REMARKS

Claims 1-41 have been cancelled pursuant to the requirement for restriction and without prejudice. Claims 43-50 have been maintained so that they can be regrouped as discussed below. New claims 51-55 have been added. The requested election of claim group has been made, and two new claim groups have proposed.

In the event that the Examiner agrees with the formation of the new claim groups, claim Group A is hereby elected. If the Examiner does not so agree, claim Group V is hereby elected. It is therefore believed that this application is in condition for examination on the merits.

The Amendments

The specification at page 49, the first paragraph of Example 7 has been amended to correct an obvious error in which grams was replaced by cubic centimeters. Enclosed Exhibit I is a page from the Handbook of Chemistry and Physics published by the Chemical Rubber Co. that lists the molecular weight for sodium meta-periodate as 213.92 and the specific gravity of the salt as 3.86 grams/cubic centimeter (ml). If 800 ml were used, that material would weigh 3088 grams [(3.86 g/ml) X (800 ml)]. That weight was said to be dissolved into 30 liters of alumina. Thus, the concentration would be 3088 g/30 L, or 102.93 g/L, which corresponds to a molarity of about 0.48. There is no support elsewhere in the application for that molarity. On the other hand, if the amended value of 800 grams were used, and those 800 g were homogeneously dispersed in the same 30 L of alumina, the concentration of meta-periodate in the alumina would be 800 g/30 L or 26.67 g/L, or [(26.67 g/213.92 g/m)/L]

0.125 molar, as value about in the middle of the range of molarities described through out the application and claims, such as claim 42 that recites a concentration of "about 0.1 to about 0.15 molar". It is thus seen that this amendment adds no new matter and simply corrects an obvious error.

The spelling of "meta-periodate" was also corrected in that same paragraph to correspond to that used in the remainder of the application and claims. See, for example the first two lines of page 14 and claim 42.

Claims 51-55 have been added. These claims are directed to an apparatus described at least in Examples 10-12. New matter has therefore not been added with these claims.

An Information Disclosure Statement, art and Form PTO-1449 were filed previously.

A fee for the added independent claim (51) and claims in addition to those already paid for that are in excess of twenty was enclosed. No further fee or petition is believed to be necessary. However, should any further fee be needed, please charge our Deposit Account No. 23-0920, and deem this paper to be the required petition. This paper is being filed in duplicate.

The Examiner is requested to phone the undersigned should any questions arise that can be dealt with over the phone to expedite this prosecution.

Respectfully submitted,

By 

Edward P. Gamson, Reg. No. 29,381

Enclosures
Exhibit I

WELSH & KATZ, LTD.
120 South Riverside Plaza, 22nd Floor
Chicago, Illinois 60606
Phone (312) 655-1500
Fax No. (312) 655-1501

CERTIFICATE OF MAILING

I hereby certify that this Election, in duplicate, along with the Petition and Fee, Added Claims fee and Exhibit I are being deposited with the United States Postal Service with sufficient postage as First Class Mail in an envelope addressed to: Commissioner for Patents, P.O. Box 1450, Alexandria, Virginia 22313-1450, on May 24, 2004.

Edward P. Gamson

EXHIBIT I

Handbook
OF
Chemistry and Physics

A Ready-Reference Book of Chemical and Physical Data

HANDBOOK OF CHEMISTRY AND PHYSICS
1974-1975
55th
EDITION
CRC PRESS, INC.

EDITOR

ROBERT C. WEAST, Ph.D.

Vice President, Research, Consolidated Natural Gas Service Company, Inc.
Formerly Professor of Chemistry at Case Institute of Technology

In collaboration with a large number of professional chemists and physicists whose assistance is acknowledged in the list of general collaborators and in connection with the particular tables or sections involved.

Published by

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18901 Cranwood Parkway • Cleveland, Ohio 44128

No.	Name	Synonyms and Formulae	Mol. wt.	Crystalline form, properties and index of refraction	Density or spec. gravity	Melting point, °C	Boiling point, °C	Solubility, in grams per 100 cc: Cold water	Hot water	Other solvents
	Sodium									d dil HCl
a289	ferrate (III)	Ferrite, $Na_2Fe_2O_4$	221.69	br, hex pl or need	4.05			18.9^{0}	67^{100}	i al
a290	ferricyanide	$Na_3Fe(CN)_6.H_2O$	298.93	red cr, deliq				31.85^{20}	156.5^{98}	i al
a291	ferrocyanide	Yellow prussiate of soda. $Na_4Fe(CN)_6.10H_2O$	484.06	pa yel, monocl, 1.519, 1.530, 1.544	1.458					i HCl; d alk
a292	fluoaluminate	Na_3AlF_6	209.94	col, monocl, 1.364	2.90	1000				
a293	fluoantimonate	$NaSbF_6$	258.73	rhomb	3.375	<1360		128.6		s al, acet
a294	fluoberyllate	Na_2BeF_4	130.99	wh, rhomb or monocl		d		1.47	2.94^{100}	
a295	fluoborate	$NaBF_4$	109.79	wh, rhomb	2.47^{20}	d 384	d	108^{26}	210^{100}	sl s al; d H_3BO_4
a296	fluoride	Nat. villiaumite. NaF	41.99	col, cub or tetr, 1.336	2.558^{41}	993	1695	4.22^{18}		s HF; i al
a297	fluoride, hydrogen	NaF.HF	61.99	col, or wh cr powd, rhdr	2.08					
a298	fluoride orthophosphate	$NaF.Na_3PO_4.12H_2O$	422.11		2.2165			12^{25}	57.5^{70}	
a299	fluoroacetate, mono-	$NaC_2H_2FO_2$	100.02	wh powd		200		111^{25}		1.4^{25} al; 5^{25} MeOH; 0.04^{25} acet; 0.0049^{25} CCl_4
a300	fluorophosphate, hexa-	$NaPF_6.H_2O$	185.97		2.369^{20}			103.2^{0}		
a301	fluorophosphate, mono-	Na_2PO_3F	143.95	col		ca 625		25		
a302	fluosilicate	Na_2SiF_6	188.06	col, hex, 1.312, 1.309	2.679	d		0.652^{17}	2.46^{100}	i al
a303	fluosulfonate	$NaSO_3F$	122.05	shiny leaf, hygr		d red heat				s al, acet; i eth
a304	formaldehyde-sulfoxylate	$NaHSO_2.CH_2O.2H_2O$	154.12	rhomb pr, hygr		64	d	v s		d a; s al, alk
a305	formate	$NaCHO_2$	68.01	col, monocl, deliq	1.92^{20}	253	d	97.2^{20}	160^{100}	sl s al; i eth
a306	2-furanacrylate	$NaC_7H_5O_3$	160.10	lt brn powd	1.919	d		s	d	sl s al; i eth
a307	metagermanate	Na_2GeO_3	166.57	wh, monocl, deliq, 1.59	3.31^{22}	1083				s a
a308	metagermanate, heptahydrate	$Na_2GeO_3.7H_2O$	292.68	col, rhomb		83		24.6		sl s al
a309	(mono-) d-glutamate	$NaC_5H_8NO_4$	169.11	wh cr		d		v s		sl s al
a310	glycerophosphate, monohydrate	$Na_2C_3H_7O_6P.H_2O$	234.05	yelsh visc liq; wh cr or powd				s		i al
a311	glycerophosphate, pentahydrate	$Na_2C_3H_7O_6P.5\frac{1}{2}H_2O$	315.12	wh pl, sc or powd		>130		v s		i al
a312	gold sulfide	$NaAuS.4H_2O$	324.08	col, monocl		d		s	d	s molten Na; i CS_2, CCl_4, NH_3, bz
a313	hydride	NaH	24.00	silver need, 1.470	0.92	d 800		d		
a314	hydroxide	NaOH	40.00	wh, deliq, 1.3576	2.130	318.4	1390	42^{0}	347^{100}	v s al, glyc; i acet, eth
a315	iodate	$NaIO_3$	197.89	wh, rhomb	$4.277^{17.5}$	d		9^{20}	34^{100}	i al; s ac a
a316	metaperiodate	$NaIO_4$	213.89	col, tetr	4.174	d 300		14.4^{25}	$38.9^{51.5}$	s H_2SO_4, HNO_3, ac a
a317	metaperiodate, trihydrate	$NaIO_4.3H_2O$	267.94	col, rhombdsh, effl	3.219	d 175		18.78	$36.4^{51.5}$	
a318	paraperiodate	Na_5IO_6	337.85	wh		800 d		d; sl s		s con NaOH sol
a319	(tri-)paraperiodate	$Na_3H_2IO_6$	293.88	col, hexag						
a320	iodide	NaI	149.89	col, cub, 1.7745	3.667^{25}	661	1304	184^{25}	302^{100}	42.57^{25} al; 39.9^{25} acet; s glyc
a321	iodide, dihydrate	$NaI.2H_2O$	185.92	col, monocl	2.448^{21}	752		317.9^{0}	1550^{100}	s NH_3; s al
a322	iodoplatinate	$Na_2PtI_6.6H_2O$	1110.59	brn, monocl	3.707			v s		s al
a323	iridium chloride	Sodium hexachloroiridate. $Na_3IrCl_6.12H_2O$	696.08	olive cr, rhomb or trig-rhomb		50		[illegible]		s al
a324	iron (III) nitrosopentacyanide	$Na_2[Fe(CN)_5NO].2H_2O$	297.95	ruby red, rhomb, 1.605, 1.575, 1.56	1.687^{20}	$-H_2O$, 100	d 160	40^{16}	s	
a325	iron (III) oxalate	$Na_3[Fe(C_2O_4)_3].5\frac{1}{2}H_2O$	487.95	grn, monocl	$1.973^{17.5}$	$-4H_2O$	d 300	32.5	182^{100}	i al
a326	iron (III) sulfate	$3Na_2SO_4.Fe_2(SO_4)_3.6H_2O$	934.09	wh, trig, 1.558, 1.613	2.5	$-6H_2O$, 100		d v sl		
a327	lactate	$NaC_3H_5O_3$	112.06	col or yelsh liq, very hygr		17	d 140	v s		s al; i eth
a328	lithium sulfate	$Na_3Li(SO_4)_2.6H_2O$	376.18	col, ditrig	2.009	$-6H_2O$, 50		s		
a329	magnesium carbonate	$Na_2CO_3.MgCO_3$	190.31	wh, rhomb	2.729^{20}	677 CO_2 1240/atm		d	d	
a330	magnesium sulfate	Nat. bloedite. $Na_2SO_4.MgSO_4.4H_2O$	334.48	col, monocl, 1.486, 1.488, 1.489	2.23			s		
a331	magnesium tartrate	$Na_2Mg(C_4H_4O_6)_2.10H_2O$	546.59	wh, monocl pr or powd				s		